

4-Chloro-2-(*p*-tolyliminomethyl)phenol

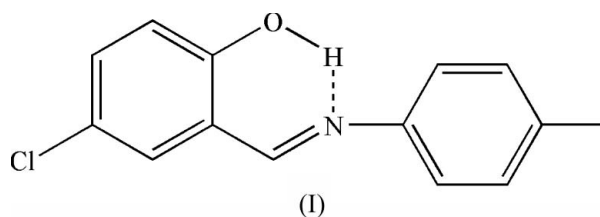
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Key indicators

Single-crystal X-ray study
 $T = 298$ K
Mean $\sigma(\text{C}-\text{C}) = 0.006$ Å
 R factor = 0.037
 wR factor = 0.099
Data-to-parameter ratio = 9.9For details of how these key indicators were
automatically derived from the article, see
<http://journals.iucr.org/e>.The title compound, $\text{C}_{14}\text{H}_{12}\text{ClNO}$, is a Schiff base compound, derived from the condensation of 5-chlorosalicylaldehyde and *p*-methoxybenzamine in MeOH. The two benzene rings are linked by a $\text{C}=\text{N}$ bond and form a dihedral angle of $44.4(2)^\circ$.Received 25 November 2006
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Comment

As an extension of our work on the structural characterization of Schiff base compounds (Li & Zhang, 2004*a,b*, 2005; Zhang & Li, 2005), the crystal structure of the title compound, (I), is reported here

In Schiff base compound (I) (Fig. 1), all the bond lengths and angles in (I) are within normal ranges (Allen *et al.*, 1987) and are comparable to those observed in a similar Schiff base compound (Kennedy & Reglinski, 2001). The two benzene rings are linked by a $\text{C}=\text{N}$ bond and form a dihedral angle of $44.4(2)$. The $\text{N1}-\text{C1}-\text{C2}-\text{C3}$, $\text{C1}-\text{N1}-\text{C2}-\text{C7}$, $\text{C1}-\text{N1}-\text{C8}-\text{C13}$ and $\text{C1}-\text{N1}-\text{C8}-\text{C9}$ torsion angles are $-6.1(6)$ and $179.8(3)^\circ$, and $149.4(3)$ and $-35.9(6)^\circ$, respectively. Atoms O1 and Cl1 deviate from the benzene ring mean plane by $-0.010(6)$ and 0.137 Å, respectively. The compound is supported by an intramolecular $\text{O}-\text{H}\cdots\text{N}$ hydrogen bond (Table 1). The $\text{C1}=\text{N1}$ bond length [$1.273(5)$ Å] confirms it to be a double bond. As expected, the molecule adopts a *trans* configuration about the $\text{C}=\text{N}$ bond.

Experimental

5-Chlorosalicylaldehyde (0.1 mmol, 15.7 mg) and *p*-toluidine (0.1 mmol, 10.7 mg) were dissolved in methanol (10 ml). The mixture was stirred at room temperature for about 30 min to give a clear yellow solution. After leaving the solution to stand in air for 12 d, yellow block-shaped crystals formed. The crystals were isolated by filtration, washed with methanol and dried in a vacuum desiccator using anhydrous CaCl_2 (yield 54%). Elemental analysis found: C 68.3, H 4.8%; calculated for $\text{C}_{14}\text{H}_{12}\text{ClNO}$: C 68.4, H 4.9%.

Crystal data

C₁₄H₁₂CINO
M_r = 245.70
 Monoclinic, *Pc*
a = 13.846 (11) Å
b = 6.986 (6) Å
c = 6.176 (5) Å
 β = 96.363 (10)°
V = 593.7 (8) Å³

Z = 2
D_x = 1.374 Mg m⁻³
 Mo *K*α radiation
 μ = 0.30 mm⁻¹
T = 298 (2) K
 Block, yellow
 0.42 × 0.25 × 0.10 mm

Data collection

Bruker SMART CCD area-detector diffractometer
 φ and ω scans
 Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
T_{min} = 0.883, *T_{max}* = 0.970

2956 measured reflections
 1531 independent reflections
 1195 reflections with *I* > 2σ(*I*)
R_{int} = 0.025
 θ_{\max} = 25.0°

Refinement

Refinement on *F*²
R [*F*² > 2σ(*F*²)] = 0.037
wR (*F*²) = 0.099
S = 1.02
 1531 reflections
 154 parameters
 H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0547P)^2 + 0.04P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$
 $\Delta\rho_{\max} = 0.15 \text{ e \AA}^{-3}$
 $\Delta\rho_{\min} = -0.21 \text{ e \AA}^{-3}$
 Absolute structure: Flack (1983),
 484 Friedel pairs
 Flack parameter: 0.000 (2)

All H atoms were placed in geometrically idealized positions and constrained to ride on their parent atoms, with C–H = 0.93–0.96 Å and *U_{iso}*(H) = 1.2*U_{eq}* or 1.5*U_{eq}*(C,O)

Data collection: *SMART* (Bruker, 2000); cell refinement: *SAINTE* (Bruker, 2000); data reduction: *SAINTE*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997*a*); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997*a*); molecular graphics: *SHELXTL* (Sheldrick, 1997*b*); software used to prepare material for publication: *SHELXTL*.

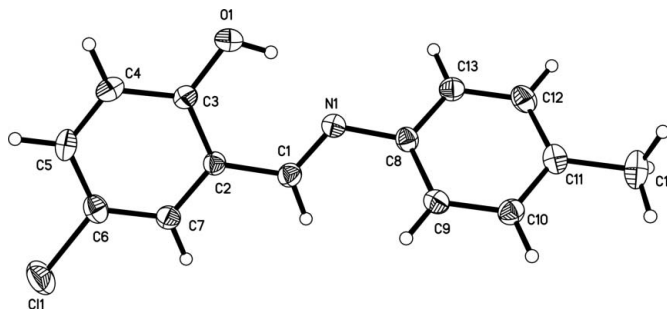


Figure 1
 The molecular structure of the title compound with 30% probability displacement ellipsoids.

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